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ब्रह्मस्य सौमनस्यं यद्वत्
तदर्थं यत्किञ्चिदनुष्ठाप्यते ।
ब्रह्मैव तेन सौमनस्यम्
अस्मै नमो न प्रीतिस्तुते ॥



एतत् सत्यं सत्यं
नृणां भद्रं यत्नादाय





घर आसं आसावं
सिध्दं प्रत्येक स्थापन पूर्ण करणं





पयः प्रशंसं स्वर्गाय

विश्वे सौमित्राणां यमुनाय





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Introduction to Management Accounting

Business

- Definition of business and profit and loss
- Internal and external stakeholders
- Business and business unit objectives
- Business strategy
- Business plan
- Business model
- Business process
- Business structure
- Business environment
- Business system
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QUESTION

- What is the difference between a **strong** and a **weak** acid?
- What is the difference between a **strong** and a **weak** base?
- What is the difference between a **strong** and a **weak** electrolyte?

ANSWER

Strong and weak acids and bases

- A **strong** acid or base is one that **completely dissociates** in water, releasing all of its **hydrogen ions** or **hydroxide ions** respectively.
- A **weak** acid or base is one that **partially dissociates** in water, releasing only a **fraction** of its **hydrogen ions** or **hydroxide ions** respectively.

- The **strength** of an acid or base is determined by its **acid dissociation constant** (K_a) or **base dissociation constant** (K_b).
- The **stronger** an acid or base, the **larger** its K_a or K_b value.

- The **strength** of an acid or base is also related to its **conjugate base** or **conjugate acid**.
- The **stronger** an acid, the **weaker** its conjugate base.

- The **stronger** a base, the **weaker** its conjugate acid.
- The **strength** of an acid or base is also related to its **structure** and **electronegativity**.

- The **stronger** an acid, the **more** electronegative the **central atom** is.
- The **stronger** a base, the **more** electronegative the **central atom** is.

QUESTION

- What is the difference between a **strong** and a **weak** acid?

ANSWER

- A **strong** acid is one that **completely dissociates** in water, releasing all of its **hydrogen ions**.
- A **weak** acid is one that **partially dissociates** in water, releasing only a **fraction** of its **hydrogen ions**.

- The **strength** of an acid is determined by its **acid dissociation constant** (K_a).
- The **stronger** an acid, the **larger** its K_a value.

- The **strength** of an acid is also related to its **structure** and **electronegativity**.
- The **stronger** an acid, the **more** electronegative the **central atom** is.

QUESTION

- What is the difference between a **strong** and a **weak** base?

ANSWER

- A **strong** base is one that **completely dissociates** in water, releasing all of its **hydroxide ions**.

QUESTION

- What is the difference between a **strong** and a **weak** electrolyte?

- A **strong** electrolyte is one that **completely dissociates** in water, releasing all of its **ions**.
- A **weak** electrolyte is one that **partially dissociates** in water, releasing only a **fraction** of its **ions**.

- The **strength** of an electrolyte is determined by its **dissociation constant** (K_d).
- The **stronger** an electrolyte, the **larger** its K_d value.

- The **strength** of an electrolyte is also related to its **structure** and **electronegativity**.
- The **stronger** an electrolyte, the **more** electronegative the **central atom** is.



